



- Q.14** The molecules of which gas has highest average kinetic energy at 25°C
- a. CO₂
 - b. O₂
 - c. CH₄
 - d. Graphite
 - e. All have same

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- Q.15 Butter melts over a wide range of temperature. On this basis, it is classified as**
a. Molecular solid
b. Crystalline solid
c. Amorphous solid
d. Covalent solid
- Q.16 Vapour pressure depends upon**
a. Amount of liquid
b. Surface area
c. Temperature
d. Shape of container
- Q.17 Van der Waal's equation is reduced to general gas equation at**
a. High temperature and low pressure
b. Low temperature and high pressure
c. High temperature and high pressure
d. Low temperature and low pressure
- Q.18 If we provide very large amount of heat to a liquid, its boiling point**
a. Remains the same
b. Increases
c. Decreases
d. Varies abnormally
- Q.19 Copper metal can be drawn into wires because**
a. Copper atoms are held together by true covalent bonds
b. Copper has unique electronic configuration
c. Copper has variable valency
d. Copper atoms are held together by non-directional metallic bonds
- Q.20 Density of gas can be determined by**
a. $\frac{PV}{RT}$
b. $\frac{PM}{RT}$
c. $\frac{nRT}{PM}$
d. PVR
- Q.21 A pair of gases with equal root mean square velocity at 300K is**
a. SO_2 , O_2
b. N_2O , CO_2
c. CO , CO_2
d. NO , NO_2
- Q.22 Which of the followings does not match?**
a. H_2O and Na^+ → Ion dipole force
b. CH_3COCH_3 and CH_3COCH_3 → dipole-dipole force
c. HCl and Ar → Dipole-dipole force
d. $\text{C}_6\text{H}_{12}\text{O}_6$ and H_2O → Hydrogen bonding
- Q.23 Which order of strength of forces is correct**
a. Dipole-Dipole Interaction > Debye forces > London forces > Hydrogen Bonding
b. Hydrogen Bonding > Dipole-Dipole Interaction > London forces > Debye forces
c. Dipole-Dipole Interaction > Hydrogen Bonding > London forces > Debye forces
d. Hydrogen Bonding > Dipole-Dipole Interaction > Debye forces > London forces
- Q.24 Pressure cooker reduces cooking time because**
a. A large flame is used
b. Boiling point of water rises
c. Heat is uniformly distributed
d. Vapour pressure of liquid reduces
- Q.25 The least volatile compound among following is**
a. H_2O
b. $\text{C}_2\text{H}_5\text{OH}$
c. NH_3
d. HF
- Q.26 _____ is NOT a property of metallic solids**
a. Malleability
b. Ductility
c. Brittleness
d. Lustrous
- Q.27 One dm^3 of each of H_2 , He , N_2 and O_2 in separate vessels at STP, have number of molecules in each.**
a. 6.02×10^{23}
b. 6.02×10^{-22}
c. 2.68×10^{22}
d. 2.68×10^{23}
- Q.28 Moving from 4°C to 0°C , the density of H_2O**
a. Increases
b. Decreases
c. 1st increase than decrease
d. Remains same
- Q.29 A gas with lowest density**
a. NH_3
b. O_2
c. Ne
d. N_2



- Q.30** Which of the following gas shows more deviation from ideal behavior at given temperature and pressure?

 - N_2
 - CO_2
 - H_2
 - He

Q.31 The Coordination number of Na^+ ion in NaCl crystal

 - 4
 - 6
 - 8
 - 12

Q.32 Which one will show malleability and ductility

 - KCl
 - Sugar
 - BN
 - Cu

Q.33 Which solid does not contain true covalent bonds?

 - Silica
 - Cadmium iodide
 - Nickel
 - Diamond

Q.34 A correct comparison of boiling point is

 - $NH_3 > HF$
 - $HF > H_2O$
 - $H_2O > HF$
 - $NH_3 > H_2O$

Q.35 During the cleansing action, the detergents attracts stain particles with a force

 - Hydrogen bonding
 - London forces
 - Dipole-induced dipole force
 - Dipole-dipole force

Q.36 Which one of the following is a solid with lowest melting point

 - NaCl
 - I_2
 - $C_6H_{12}O_6$
 - Fe

Q.37 Mathematically Boyle's law is shown by all except

 - $PV = K$
 - $PT = K$
 - $P_1V_1 = P_2V_2$
 - $\frac{V_1}{V_2} = \frac{P_2}{P_1}$

Q.38 At $100^\circ C$ a gas has 1 atm pressure and $10dm^3$ volume, its volume at STP would be

 - $10dm^3$
 - More than $10dm^3$
 - Less than $10dm^3$
 - Can't be predicted

Q.39 By Charle's law, there will be a change in the volume of a given mass of gas by $1/273$ of its original volume at $0^\circ C$, if the temperature of gas is changed by

 - $10^\circ C$
 - $1^\circ C$
 - $100^\circ C$
 - $2^\circ C$

Q.40 When we plot a graph between pressure on X-axis and the product PV on Y-axis. A straight line parallel to the pressure axis is obtained. This straight line

 - Is called isotherm
 - Will help us to understand the non-ideal behaviour of gases
 - Can change its position by changing temperature
 - All of these are correct

Q.41 A real gas under what conditions will behave non ideally

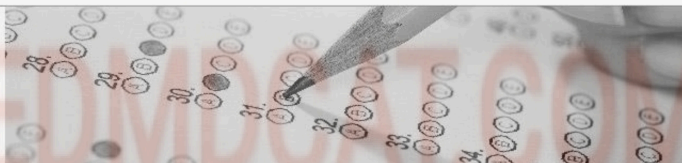
	Temperature	Pressure
a.	Low	Low
c.	High	High
b.	High	Low
d.	Low	High

Q.42 H_2O is liquid at room temperature whereas H_2S is a gas because

 - H_2O is used for drinking, but H_2S has rotten egg smell
 - H_2O is neutral, H_2S is weak acid
 - H_2O is more abundant than H_2S
 - H_2O has hydrogen bonding but H_2S has no hydrogen bonding

Q.43 Boiling point of H_2O is higher than that of HF although fluorine is more electronegative than oxygen. It is due to

 - Atomic radius of fluorine is bigger than oxygen
 - H_2O is neutral HF is acidic nature
 - Geometry of H_2O is angular, but HF is linear
 - H_2O forms two hydrogen bonds per molecule but HF forms one



- Q.44** Vapour pressure of liquid is measured when liquid and the vapours are in equilibrium it means that
- Liquid and vapours have same value of kinetic energy
 - Liquid and vapours have same heat content
 - Rate of evaporation is equal to the rate of condensation
 - Rate of evaporation and condensation are different
- Q.45** The strongest H-bond is
- $\text{H}-\text{O}^{\delta-} \cdots \text{H}^{\delta+}-\text{O}$
 - $\text{H}-\text{F}^{\delta-} \cdots \text{H}^{\delta+}-\text{F}$
 - $\text{H}-\text{N}^{\delta-} \cdots \text{H}^{\delta+}-\text{N}$
 - $\text{H}-\text{Cl}^{\delta-} \cdots \text{H}^{\delta+}-\text{Cl}$
- Q.46** The density of gas 'X' is twice than that of 'Y' under same conditions. If molar mass of 'X' is 'M' then the molar mass of 'Y' is
- M/2
 - 2M
 - M
 - 4M
- Q.47** Water boils at 25°C if external pressure is
- 323 torr
 - 700 torr
 - 23.7 torr
 - 1489 torr
- Q.48** An example of covalent solid which has three dimensional network structure is
- Ice
 - Diamond
 - Sodium chloride
 - Graphite
- Q.49** The chloroform and acetone are miscible due to hydrogen bonding. The type of force between chloroform molecules is
- Hydrogen bonding
 - Dipole – dipole force
 - Van der Waals's force
 - Dipole induced dipole force
- Q.50** If 20 g of a gas at 1 atmosphere pressure is cooled from 273°C to 0°C at constant volume its pressure would become
- 0.25 atm
 - 1.5 atm
 - 1.0 atm
 - 0.5 atm

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Chemistry

			21 B	31 B	41 D
C	11 B				
			22 C	32 D	42 D
A	12 B			33 C	43 D
			23 D		
C	13 B			34 C	44 B
			24 B		
B	14 D			35 B	45 B
			25 A		
D	15 C			36 B	46 A
			26 C		
D	16 C			37 B	47 C
C	17 A		27 C		
			28 B	38 C	48 B
C	18 A				
			29 A	39 B	49 B
C	19 D				
			30 B	40 D	50 D
A	20 B				

Physics

D	11 B	21 B	31 B	41 D
B	12 C	22 D	32 A	42 B
B	13 D	23 B	33 D	43 C
B	14 C	24 B	34 D	44 B
B	15 C	25 B	35 C	45 C
C	16 A	26 B	36 C	46 C
C	17 B	27 D	37 B	47 A
10 C	18 B	28 D	38 B	48 A
D	19 A	29 B	39 B	49 D
B	20 A	30 B	40 A	50 B

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